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Molecular, complete and net ionic equations worksheet

Write the balanced molecular, complete ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in aqueous solution. If no reaction occurs, write "no reaction".



Atoms and ions worksheet

Name \_\_\_\_\_ Period \_\_\_\_\_

Fill in the missing values in the table below.

Element	Symbol	Atomic Number	Mass Number	Protons	Neutrons	Electrons
Carbon		6	12			
	$\text{C}^{2-}$					8
Polonium		84	210	84	126	
	${}_{84}^{\text{Po}}$					
Tin	$\text{Sn}^{2+}$	50				48
Zinc			64	30		
			66			30
			68	30		
Cobalt	$\text{Co}^{3+}$	27				24
Iron			56	26	30	26
	$[\text{Fe}]$					
		26			30	

Ion	Ion Symbol	Atomic Number	Mass Number	Protons	Neutrons	Electrons	Overall Atomic Charge
Sulfide	$\text{S}^{2-}$		32	16		18	
	$\text{Cl}^{-}$		35	17			
Hydrogen		1	1	1		1	
	$\text{H}^{+}$	1			0		
		1	1			1	
	$\text{H}^{-}$	1	1			2	
Oxygen	$\text{O}^{2-}$	8	16	8		10	
Nitrogen			14	7	7	7	
Aluminum	$\text{Al}^{3+}$	13	27	13		10	
	$\text{Fe}^{2+}$	26	56	26		24	
	$\text{Ca}^{2+}$	20	40	20	20	18	
		20	40	20	20	20	

Net Ionic Equations

1.  $\text{Fe}^{3+}(\text{aq}) + 3\text{OH}^{-}(\text{aq}) \rightarrow \text{Fe}(\text{OH})_3(\text{s})$
2.  $\text{CO}_3^{2-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$
3.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
4.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
5.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
6.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
7.  $\text{CO}_3^{2-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$
8.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
9.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
10.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
11.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
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18.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
19.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
20.  $2\text{OH}^{-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$



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